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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C₃H₄(g) + x. O₂(g) (3CO₂(g) + 2H₂O(g) a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C₃H₄? c. How many moles are in an 8.0 g sample of C₃H₄? 2. a. What is meant by . ideal conditions

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Stoichiometry b. Theoretically, how many moles of NH₃ will be produced? PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the ...

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